Balancing Chemical Equations Practice

Name _____

Balancing Equations Practice

Write a coefficient on each blank so the equation is balanced (write 1 if only one is needed). Then fill in the table with the atom counts after balancing.

1 ___ N_2 + ___ H_2 → ___ NH_3

Atom	Reactants	Products
N		
H		

2 ___ Na + ___ Cl_2 → ___ NaCl

Atom	Reactants	Products
Na		
Cl		

3 ___ Li + ___ O_2 → ___ Li_2O

Atom	Reactants	Products
Li		
O		

4 ___ H_2 + ___ Cl_2 → ___ HCl

Atom	Reactants	Products
H		
Cl		

5 ___ $KClO_3$ → ___ KCl + ___ O_2

Atom	Reactants	Products
K		
Cl		
O		

6 ___ Al + ___ Cl_2 → ___ $AlCl_3$

Atom	Reactants	Products
Al		
Cl		

Balancing Equations Practice Answer Key

Answers are in green. A coefficient of 1 (grey) is usually left blank.

1 $_1_\ N_2 + _3_\ H_2 \rightarrow _2_\ NH_3$

Atom	Reactants	Products
N	2	2
H	6	6

2 $_2_\ Na + _1_\ Cl_2 \rightarrow _2_\ NaCl$

Atom	Reactants	Products
Na	2	2
Cl	2	2

3 $_4_\ Li + _1_\ O_2 \rightarrow _2_\ Li_2O$

Atom	Reactants	Products
Li	4	4
O	2	2

4 $_1_\ H_2 + _1_\ Cl_2 \rightarrow _2_\ HCl$

Atom	Reactants	Products
H	2	2
Cl	2	2

5 $_2_\ KClO_3 \rightarrow _2_\ KCl + _3_\ O_2$

Atom	Reactants	Products
K	2	2
Cl	2	2
O	6	6

6 $_2_\ Al + _3_\ Cl_2 \rightarrow _2_\ AlCl_3$

Atom	Reactants	Products
Al	2	2
Cl	6	6